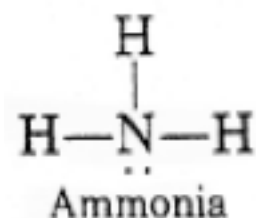


Pharmaceutical Organic Chemistry-1

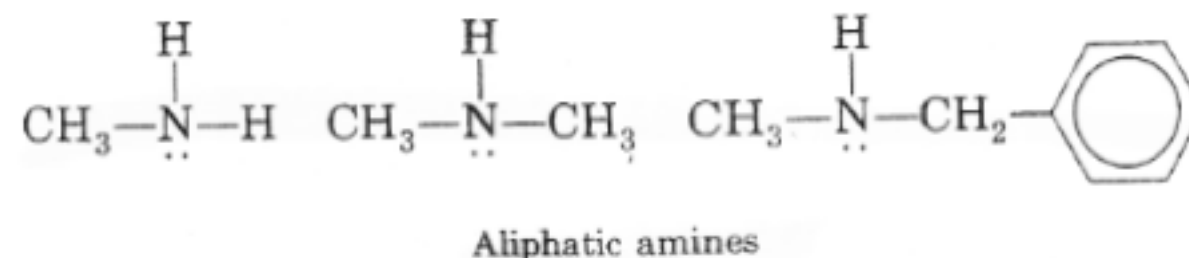
Chapter-8: Amines

Structure and Classification of Amines

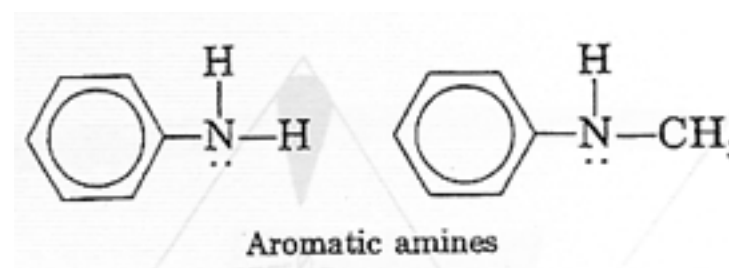
- **Amines** are compounds that derived from **ammonia** by replacement of one, two, or three hydrogens by alkyl or aryl groups.



- **Aliphatic amines** contain *only alkyl* groups bonded directly to the nitrogen atom.

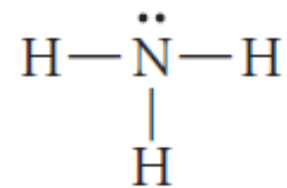


- **Aromatic amines** are those in which one or more aryl groups are bonded directly to nitrogen.

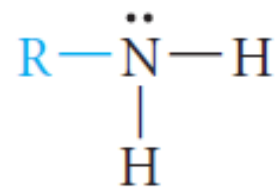


Classification and Structure of Amines

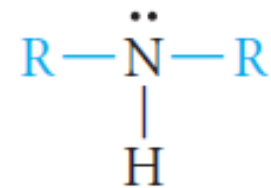
- The relation between **ammonia and amines** is illustrated by the following structures:



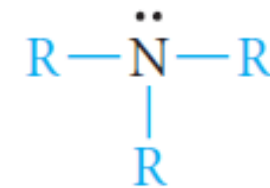
ammonia



primary amine



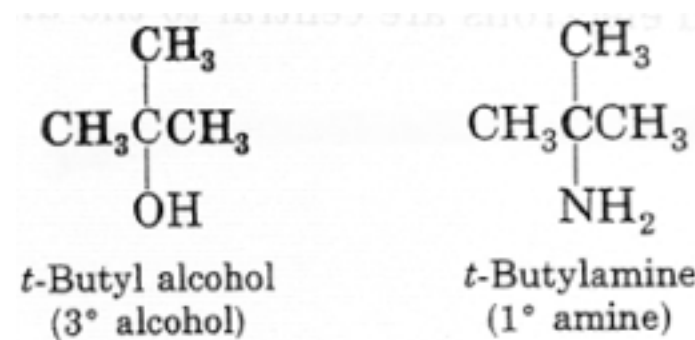
secondary amine



tertiary amine

- Amines** are classified as **primary**, **secondary**, or **tertiary**, *depending on whether one, two, or three organic groups are attached to the nitrogen.*

- NOTE:**



***t*-butyl alcohol** is a **tertiary alcohol** (because three carbons are attached to the carbinol carbon).

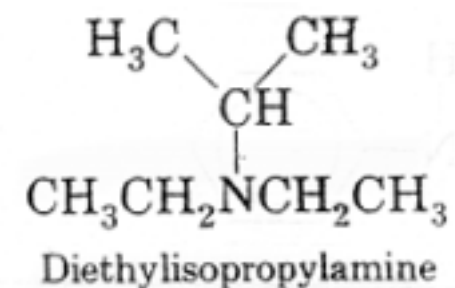
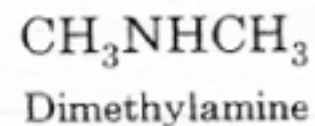
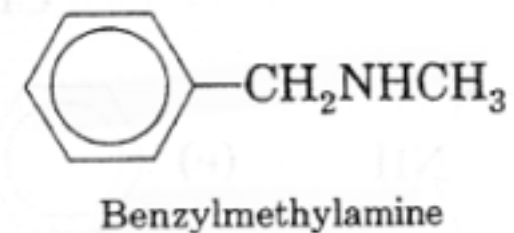
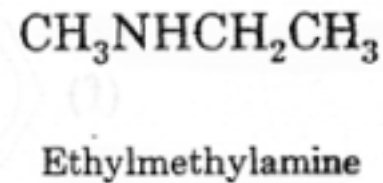
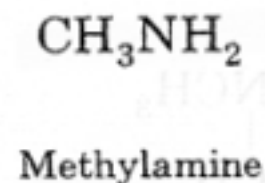
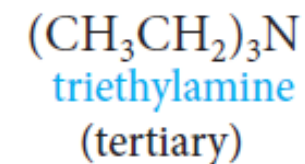
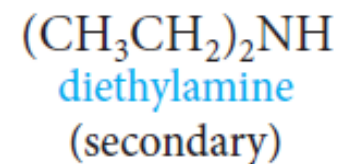
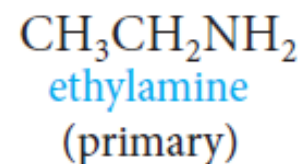
***t*-butyl amine** is a **primary amine** (because only one carbon is attached directly to the nitrogen atom).

Nomenclature of Amines

Common

Names

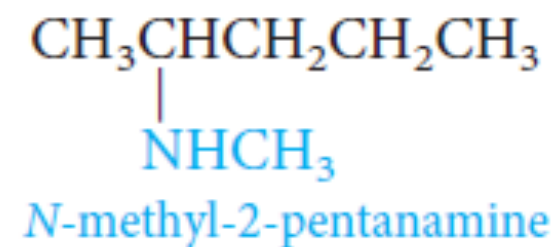
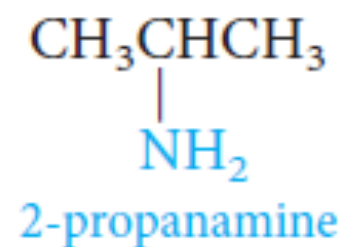
Amines are named by specifying the alkyl groups attached to the nitrogen and adding the suffix *-amine* (*Alkylamine*).



Nomenclature of Amines

I U P A C System

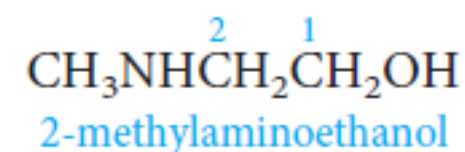
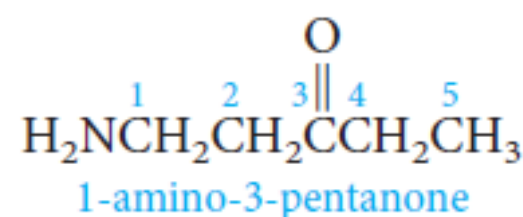
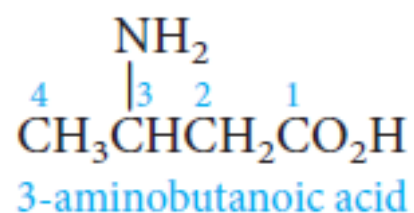
- Amines can be named as **alkanamines**.



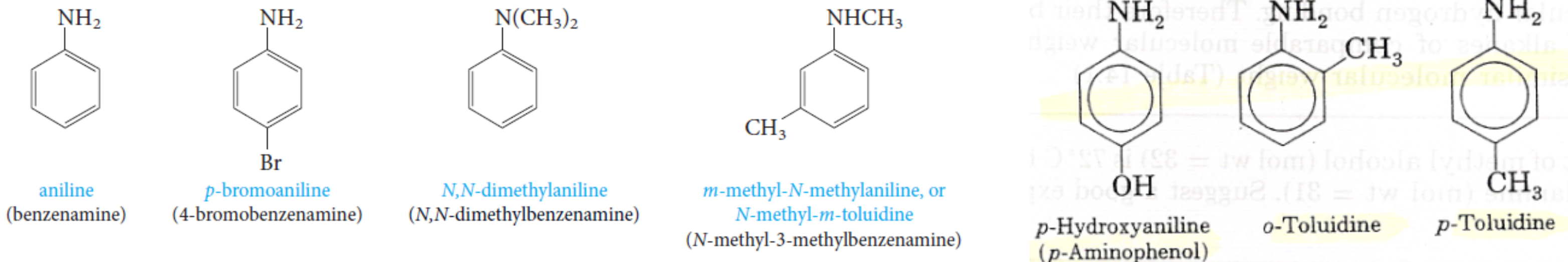
Nomenclature of Amines

IUPAC

System When **other functional groups** are present, the amino group, -NH_2 , is named as a substituent.



- **Aromatic amines** are named as derivatives of aniline.
- In the IUPAC system, aniline is called benzenamine.



Physical Properties of Amines

Boiling

- **Methylamine and ethylamine are gases**, but primary amines with three or more carbons are liquids.
- **Primary amines** boil well above alkanes with comparable molecular weights, but below comparable alcohols.

Intermolecular N-H · · · N hydrogen bonds are important and raise the boiling points of primary and secondary amines but are not as strong as the O-H · · · O bonds of alcohols.

The reason for this is that nitrogen is not as electronegative as oxygen

alkane	CH ₃ CH ₃ (30) bp −88.6°C	CH ₃ CH ₂ CH ₃ (44) bp −42.1°C
amine	CH ₃ NH ₂ (31) bp −6.3°C	CH ₃ CH ₂ NH ₂ (45) bp +16.6°C
alcohol	CH ₃ OH (32) bp +65.0°C	CH ₃ CH ₂ OH (46) bp +78.5°C

Physical Properties of Amines

Boiling

Tertiary amines are also polar compounds, but because hydrogen is not bonded to nitrogen, these amines are incapable of **intermolecular hydrogen bonding**.

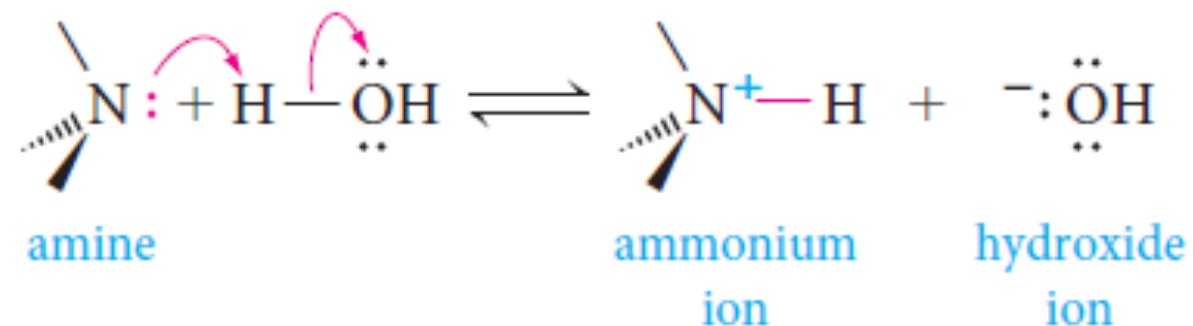
Their boiling points are Lower than primary and secondary amines of identical molecular weights and Higher than those of alkanes of similar molecular weight.

Solubility in

- **All three classes of amines** can form hydrogen bonds with the -OH group of water (that is, $\text{O}-\text{H} \cdots \text{N}$).
- **Primary and secondary amines** can also form hydrogen bonds with the oxygen atom in water: $\text{N}-\text{H} \cdots \text{O}$.
- **Amines** with up to six carbons show appreciable solubility in water.

The Basicity of Amines

- The **unshared pair of electrons** on the nitrogen atom dominates the chemistry of amines.
- Because of this electron pair, **amines are both basic and nucleophilic**.
- Aqueous solutions of amines are basic because of the following equilibrium:



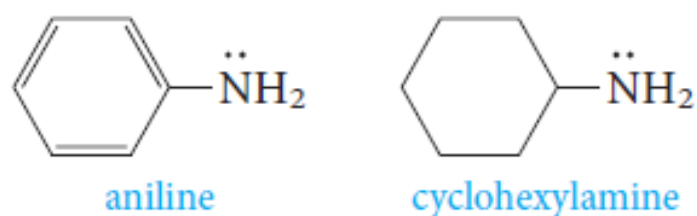
☐ Electron-donating groups increase the basicity of amines.

☐ Electron-withdrawing groups decrease their basicity

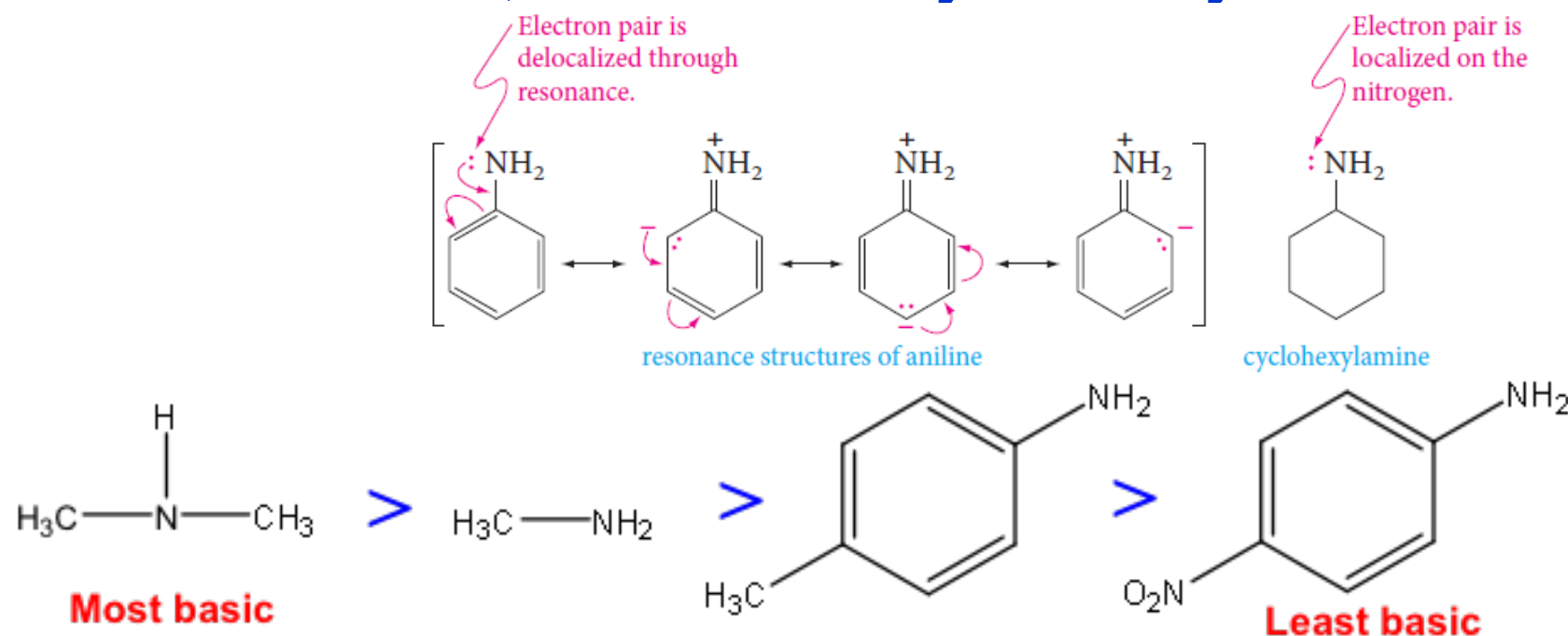


The Basicity of Amine

- **Aromatic amines are much weaker than aliphatic amines or ammonia.**
 - **Example:** aniline is less basic than cyclohexylamine.



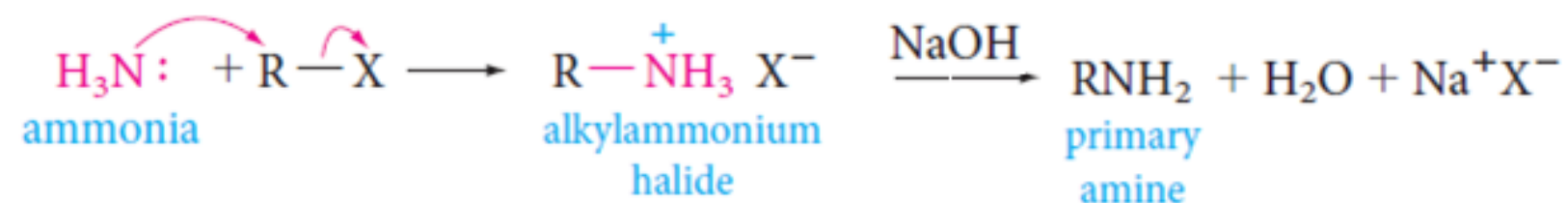
The reason is the resonance delocalization of the unshared electron pair that is possible in aniline, but not in cyclohexylamine:



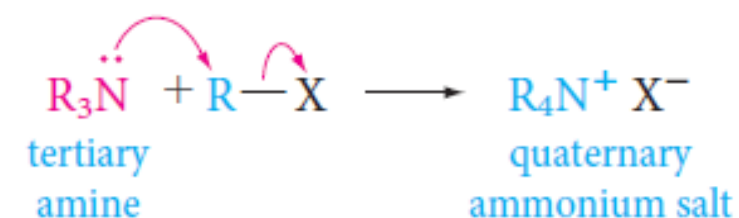
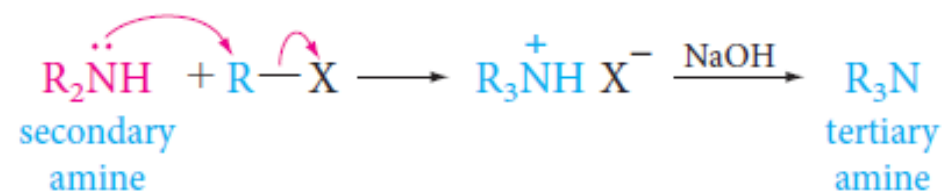
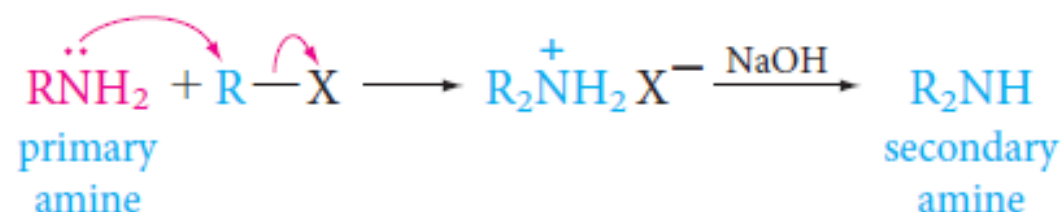
Preparation of Amines

1) Alkylation of Ammonia

- **Ammonia** reacts with alkyl halides to give amines via a two-step process.
The first step is a nucleophilic substitution reaction. The free amine can then be obtained from its salt by treatment with a strong base



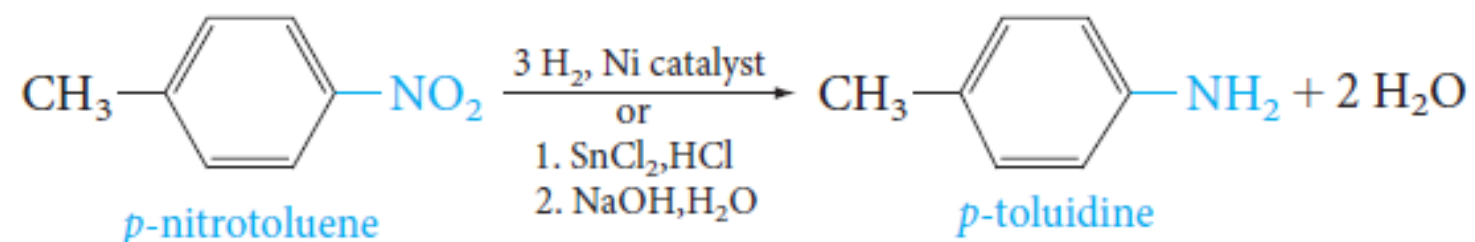
- **Primary, secondary, and tertiary amines** can be similarly alkylated.



2) Reduction of Nitro Groups

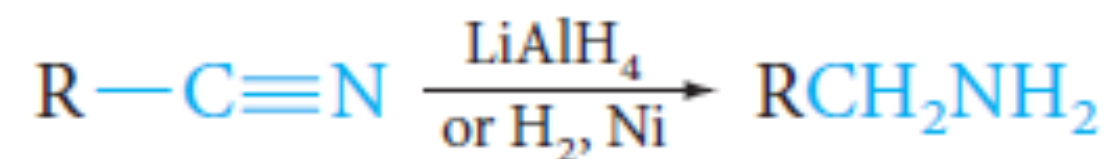
- The best route to **aromatic primary amines** is by reduction of the corresponding nitro compounds.

The nitro group is easily reduced, either catalytically with hydrogen or by chemical reducing agents.



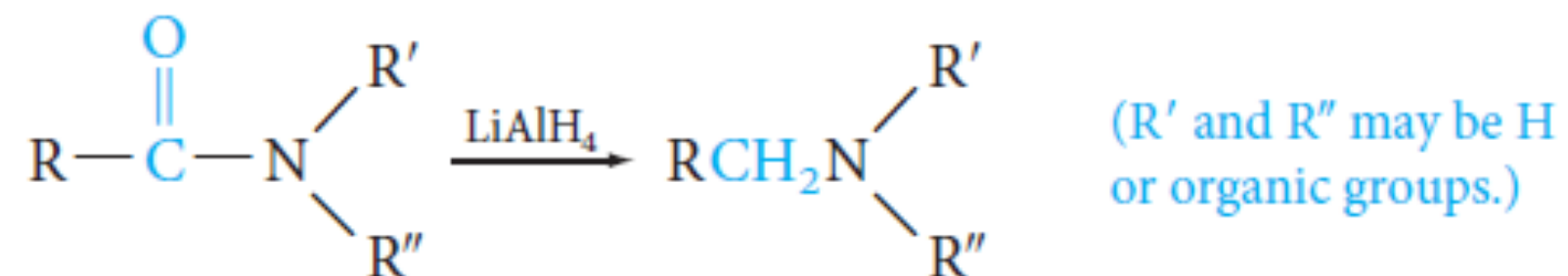
3) Reduction of Nitriles

- **Reduction of nitriles** (cyanides) gives primary amines.



4) Reduction of Amides

- **Amides** can be reduced to amines with lithium aluminum hydride.

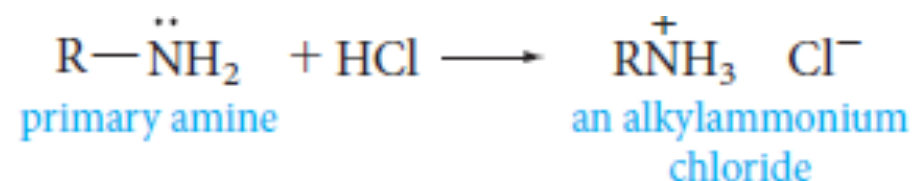


Reactions of Amines

1) Reactions with Acids: Salt

Formation

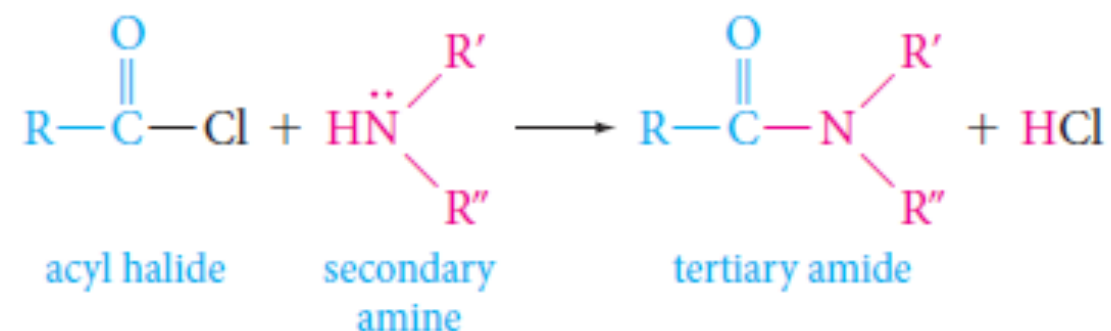
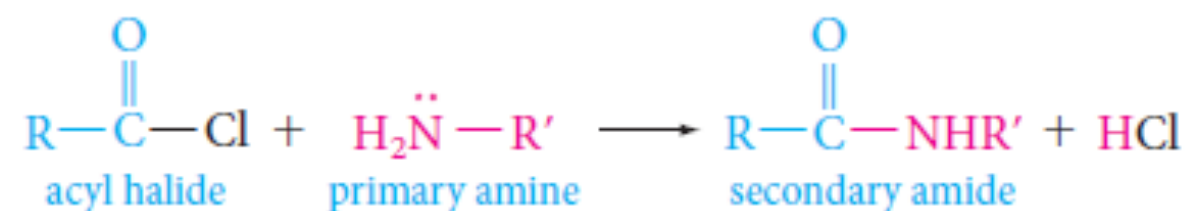
Amines react with strong acids to form alkylammonium salts.



2) Acylation of Amines: Amides

Formation

Primary and secondary amines react with acyl halides to form amides.



3) Imines

Formation Primary amines, $R-NH_2$ or $ArNH_2$, undergo nucleophilic addition with aldehydes or ketones in an acidic buffer to give substituted imines.

